The Instructor reserves the right to alter or amend this syllabus to meet instructional goals.

Instructor: Christopher D. Kim, Ph.D.
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E-Mail Address: cdk_delta@yahoo.com
ckim@deltacollege.edu
Phone: (209) 954-5343
Office Hours: MWF 7:00 am to 8:00 am
TTh 1:00 pm to 2:00 pm
Lecture Room: Cunningham 333
MWF 8:00 am to 9:00 am
Lab Room: Cunningham 301
MW 9:00 am to 11:00 am (16893)
MW 12:00 pm to 2:00 pm (83073)
Website: http://www.deltacollege.edu/emp/ckim

Prerequisites: Reading Level Two, Chemistry 3A, and Math 82, both with a grade of “C” or better.

College Statement: This course is designed as a technical introduction to chemistry for students intending to major in chemistry, medicine, dentistry, pharmacy, etc., which require a rigorous understanding of the fundamental principles of chemistry.

General Aims: Upon successful completion of this course, the student will be able to demonstrate an understanding of the principles of general chemistry and apply them to different scientific fields of study and laboratory work.

Specific Objectives: The student will be able to:
1. Demonstrate the ability to apply dimensional analysis to chemical problems.
2. Demonstrate theoretical principles in atomic structure and how they relate to the Periodic Table and thermochemistry including heats of formation and applications.
3. Demonstrate an understanding of inorganic nomenclature, chemical reactions, and molar relationships.
4. Describe and apply the physical laws of gases in both chemical and biological systems.
5. Describe, discuss and apply theoretical principles concerning the physical states of matter.
6. Demonstrate an understanding of the nature of chemical bonding.
7. Demonstrate an understanding of introductory organic nomenclature and reactions with emphasis upon functional groups related to both chemistry and biology.
8. Demonstrate theoretical principles in solution chemistry with emphasis upon concentration units, dilutions, and chemical reactions.
9. Demonstrate a basic understanding of laboratory procedure and chemical safety.
10. Demonstrate skills in analytical and physical techniques in the laboratory.
Specific Objectives: 11. Demonstrate skills in observation and interpretation of chemical changes.

Materials

Required Materials:
- Chemistry by Zumdahl & Zumdahl, current edition
- Chemical Principles in the Laboratory by Slowinski, Wolsey, & Masteron, current edition
- Scientific Calculator (Storage calculators will NOT be allowed during quizzes or exams.)
- Periodic Table (printable version from http://www.webelements.com)
- Lab Goggles
- Lock

Grading

The four hourly exams will be worth 100 points each. The comprehensive final exam will be worth 200 points.

<table>
<thead>
<tr>
<th>Exam</th>
<th>Textbook Chapters</th>
</tr>
</thead>
<tbody>
<tr>
<td>1</td>
<td>1, 2.1 – 2.7, 7</td>
</tr>
<tr>
<td>2</td>
<td>2.8, 3, 4, 5</td>
</tr>
<tr>
<td>3</td>
<td>6, 8, 9</td>
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<tr>
<td>4</td>
<td>10, 11</td>
</tr>
<tr>
<td>Final</td>
<td>1-11 &amp; Organic Nomen.</td>
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</tbody>
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Quizzes will be worth 10 points each. Approximately 11 quizzes will be given. The two lowest quizzes will be dropped giving a total of 100 quiz points.

There will be approximately 11 laboratory exercises will be performed. Each lab will be worth 20 point each. A 50-point lab exam will be given.

APPROXIMATE point breakdown (subject to change):

<table>
<thead>
<tr>
<th></th>
<th>Points</th>
<th>Percentage</th>
</tr>
</thead>
<tbody>
<tr>
<td>Exams</td>
<td>400</td>
<td>(~41.2% of your grade)</td>
</tr>
<tr>
<td>Final Exam</td>
<td>200</td>
<td>(~20.6% of your grade)</td>
</tr>
<tr>
<td>Quizzes</td>
<td>100</td>
<td>(~10.3% of your grade)</td>
</tr>
<tr>
<td>Lab Exercises</td>
<td>220</td>
<td>(~22.7% of your grade)</td>
</tr>
<tr>
<td>Lab Exam</td>
<td>50</td>
<td>(~5.2% of your grade)</td>
</tr>
</tbody>
</table>

970 points

Grade assignments will be made based on the following scale:

<table>
<thead>
<tr>
<th>Grade</th>
<th>Percentage</th>
</tr>
</thead>
<tbody>
<tr>
<td>A</td>
<td>90.0 % - 100 %</td>
</tr>
<tr>
<td>B</td>
<td>78.0 % - 89.9 %</td>
</tr>
<tr>
<td>C</td>
<td>65.0 % - 77.9 %</td>
</tr>
<tr>
<td>D</td>
<td>50.0 % - 64.9 %</td>
</tr>
<tr>
<td>F</td>
<td>0 % - 49.9 %</td>
</tr>
</tbody>
</table>
CLASSROOM and LABORATORY expectations:

- All students are expected to RESPECT themselves, one another, the instructor, the room, and the equipment. In turn, I will respect students and their academic needs and progress.

- **REGULAR ATTENDANCE** to lecture and laboratory periods is expected. Time lost to tardiness to lab, quizzes, or exams cannot be made up. Other classes are scheduled into our lab and lecture rooms immediately after our class ends. Exams, quizzes, and lab exercises may not be made-up. If an unavoidable conflict exists, for example a death in the family, arrangements may be made. Keep in touch with me if a problem arises. Use email or use the phone!

- **ACADEMIC INTEGRITY**: *Cheating or academic dishonesty of any kind will not be tolerated!*
  The FIRST offense will result in the most severe consequences as outlined in the Student Handbook. The FIRST offense will result in a grade of zero on the item in question (will NOT be dropped). The SECOND offense will result in course failure. Please see the Student Handbook or Course Catalog for the college's definition of academic dishonesty and its consequences. Please be aware that I may be implementing a new system to deter cheating.

- **SAFETY**: All students are expected to abide by the safety rules in the laboratory. Contact lens are not to be worn in lab. Note that safety glasses or goggles are required at all times in the laboratory.

- **Special Needs**
  If you have any special needs for accessibility or any other issues (ex: asthma or pregnancy) please discuss with me so that appropriate accommodations may be made.

- Please SILENCE mobile phones and pagers before entering the lab or classroom.

- **Other important dates**: Last drop date without a W: 2/6/2004
  Last drop date with a W: 4/22/2004

**IT IS THE STUDENT’S RESPONSIBILITY TO DROP THE CLASS**