

CHEMISTRY 3A --- WORKSHEET 8

CHAPTER 5

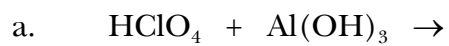
1. State whether the following solutions would be acidic, basic, or neutral.
 - a. Ivory Liquid
 - b. $[\text{H}^+] = 3.0 \times 10^{-7}$
 - c. $[\text{OH}^-] = 0.50 [\text{H}^+]$
 - d. $\text{pOH} = 7.15$

2. Identify the acid, base, conjugate acid, and conjugate base in each of the following equations.
 - a. $\text{HF} + \text{NH}_3 \leftrightarrow \text{NH}_4^+ + \text{F}^-$

 - b. $\text{H}_2\text{PO}_4^- + \text{HS}^- \leftrightarrow \text{HPO}_4^{2-} + \text{H}_2\text{S}$

3. Identify the following as strong acids, weak acids, strong bases, or weak bases.
 - a. HBr
 - b. HCN
 - c. CsOH
 - d. NH_3

4. For each of the following reactions, show the products that would be expected to form and balance the equation. Be sure to label the state of each substance, e.g. (s), (aq), etc.



5. Write the name or formulas for the following acids or bases.

