

NAMES, FORMULAS AND CHARGES OF COMMON IONS

	Positive Ions (Cations)			Negative Ions (Anions)			
1+	Ammonium	NH_4^+	1-	Acetate	$\text{C}_2\text{H}_3\text{O}_2^-$		
	Copper(I) (Cuprous)	Cu^+		Bromate	BrO_3^-		
	Hydrogen	H^+		Bromide	Br^-		
	Potassium	K^+		Chlorate	ClO_3^-		
	Silver	Ag^+		Chloride	Cl^-		
	Sodium	Na^+		Chlorite	ClO_2^-		
2+	Barium	Ba^{2+}	2-	Cyanide	CN^-		
	Cadmium	Cd^{2+}		Fluoride	F^-		
	Calcium	Ca^{2+}		Hydride	H^-		
	Cobalt(II)	Co^{2+}		Hydrogen carbonate (Bicarbonate)	HCO_3^-		
	Copper(II) (Cupric)	Cu^{2+}		Hydrogen sulfate (Bisulfate)	HSO_4^-		
	Iron(II) (Ferrous)	Fe^{2+}		Hydrogen sulfite (Bisulfite)	HSO_3^-		
	Lead(II)	Pb^{2+}		Hydroxide	OH^-		
	Magnesium	Mg^{2+}		Hypochlorite	ClO^-		
	Manganese(II)	Mn^{2+}		Iodate	IO_3^-		
	Mercury(II) (Mercuric)	Hg^{2+}		Iodide	I^-		
	Nickel(II)	Ni^{2+}		Nitrate	NO_3^-		
	Tin(II) (Stannous)	Sn^{2+}		Nitrite	NO_2^-		
	Zinc	Zn^{2+}		Perchlorate	ClO_4^-		
	3+	Aluminum		Al^{3+}	2-	Permanganate	MnO_4^-
Antimony(III)		Sb^{3+}	Thiocyanate	SCN^-			
Arsenic(III)		As^{3+}	Carbonate	CO_3^{2-}			
Bismuth(III)		Bi^{3+}	Chromate	CrO_4^{2-}			
Chromium(III)		Cr^{3+}	Dichromate	$\text{Cr}_2\text{O}_7^{2-}$			
Iron(III) (Ferric)		Fe^{3+}	Oxalate	$\text{C}_2\text{O}_4^{2-}$			
Titanium(III) (Titanous)		Ti^{3+}	Oxide	O^{2-}			
4+		Manganese(IV)	Mn^{4+}	3-		Peroxide	O_2^{2-}
		Tin(IV) (Stannic)	Sn^{4+}			Silicate	SiO_3^{2-}
	Titanium(IV) (Titanic)	Ti^{4+}	Sulfate		SO_4^{2-}		
5+	Antimony(V)	Sb^{5+}	3-	Sulfide	S^{2-}		
	Arsenic(V)	As^{5+}		Sulfite	SO_3^{2-}		
3-			3-	Arsenate	AsO_4^{3-}		
				Borate	BO_3^{3-}		
				Phosphate	PO_4^{3-}		
				Phosphide	P^{3-}		
				Phosphite	PO_3^{3-}		