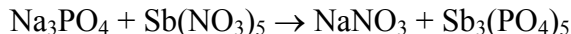
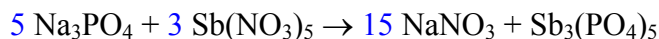


Writing the Net Ionic Reaction

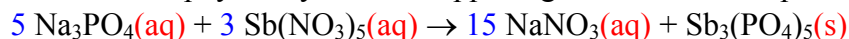


Example Reaction

- 1) Step one BALANCE the overall chemical reaction.



- 2) Use the PPT Table or physically what is happening to determine the phases.



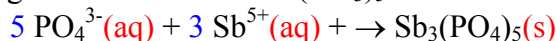
- 3) Bring down whatever compound is a solid, liquid, gas, or weak electrolyte in to the net ionic reaction.



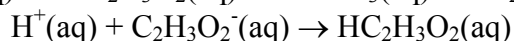
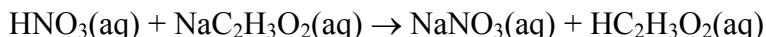
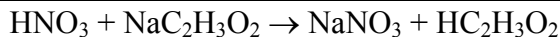
Net Ionic Reaction

- 4) Focus on the products of the reaction.

- Remember if the phase is (aq) it means that the compound is dissolved in water and can be broken up in to its ions (as long as it is ionic and not a weak electrolyte).
- Think about how the solid, liquid, gas, or weak electrolyte is made on the product side. Since Na_3PO_4 is (aq), ionic, and not a weak electrolyte it breaks up into its ions; the PO_4^{3-} comes down into the net ionic reaction to make the solid.
- The same argument is made for $\text{Sb}(\text{NO}_3)_5$ and the use of Sb^{5+} .



EXAMPLES

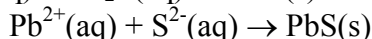
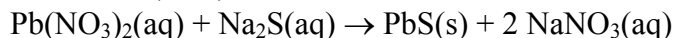
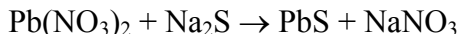


(It really looks like there is no reaction occurring because all the compounds are aqueous, but since $\text{HC}_2\text{H}_3\text{O}_2$ is a weak electrolyte a reaction occurs.)

Overall Reaction

Balanced Reaction & States

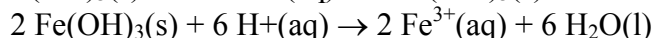
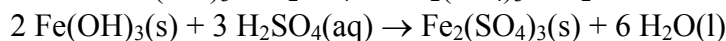
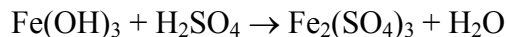
Net Ionic Reaction



Overall Reaction

Balanced Reaction & States

Net Ionic Reaction



Overall Reaction

Balanced Reaction & States

Net Ionic Reaction