

Exam 2 Topics

- **Chapter 2.7**
 1. Ions!
 - Monatomic
 - Polyatomic
 2. Nomenclature
 - Ionic Compounds
 - Non-metal Binary Compounds
 - Acids
- **Chapter 3**
 1. BALANCED EQUATIONS
 2. Conservation of Mass
 3. Limiting Reactant
 4. Actual Yield / Theoretical Yield / % Yield
 5. MOLES
 6. Molar mass
 7. Empirical Formula / Molecular Formula calculations
 8. Hydrated Compounds
 9. % Composition
- **Chapter 4**
 1. Electrolytes & non-electrolytes
 2. Solubility table
 3. 3 types of aqueous reactions
 - Precipitation
 - Acid/Base
 - Neutralization
 - Gas forming
 - REDOX (Balance)
 4. Net ionic equations
 5. Oxidation numbers
 6. Molarity
 7. pH
 - $\text{pH} = -\log [\text{H}^+]$
- **Chapter 5**
 1. Boyle's Law
 2. Charles's Law
 3. General gas law
 4. Ideal gas law
 - Density
 5. Partial pressure
 6. Rate of effusion & diffusion
 7. Kinetic molecular theory (Collision theory)